



CENTRE FOR VOCATIONAL STUDIES

SCHOOL OF HEALTH SCIENCES

ISLAMIC UNIVERSITY OF SCIENCE AND TECHNOLOGY, KASHMIR

Admission to BVoc in Medical Laboratory and Molecular Diagnostic Technology

ELIGIBILITY:

Candidates seeking admission to this program should have passed 10+2 with a minimum of 45% marks from the J&K Board of School Education or from any other recognized board with Physics, Chemistry & Biology as compulsory subjects.

MODE OF SELECTION:

The selection will be based on candidate's performance in the University Entrance Test and percentage of marks in aggregate obtained by a candidate in the qualifying examination with a ratio of 60:40 respectively. The syllabus for the entrance exam consists of Biology, Physics and Chemistry subjects. Students can refer to their Class 11th and 12th syllabus for all these subjects and prepare for the entrance exam, to get admission in this Course.

INTAKE CAPACITY: 30

DURATION: 03 years (06 semesters)



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Syllabus for the Entrance Test



PHYSICS

25% of the maximum marks is allotted to numerical problems.

Maximum Marks: 100

Maximum Marks: 100
Theory: 70 Marks

Time: 3 hours
Practical: 30 Marks

Unit - I : Mathematical Tools

Marks 04

Functions, limits of function, simple ideas of differentiation integration, differentiation of x^n , e^{ax} , $\sin x$ by ab-initio method, integration of x^n , $1/x$, e^{ax} , $\sin x$ and $\cos x$. Simple Idea of definite integral.

Unit - II : Physical world and measurement

Marks 5

Physics - Scope and excitement, physics in relation to science, society and technology. Need for measurement, units of measurement, system of units, SI Units, fundamental and derived units, length, mass and time measurement. Accuracy and precision of measuring instruments; errors in measurement, significant figures.

Dimensions of physical quantities, dimensional analysis, its applications.

Unit III : Kinematics

Marks 7

Motion in a straight line, position time graph, speed and velocity.

Uniform and non uniform motion, average speed and instantaneous velocity. Uniformly accelerated motion, velocity-time graph, position time graphs, relations for uniformly accelerated motion. (graphical treatment and calculus approach).

Scalar and vector quantities, position and displacement vectors, general vector and notation, equality of vectors, multiplication of vectors by a real number, addition and subtraction of vectors, Relative velocity.

Unit vector, Resolution of a vector in a plane rectangular components, Scalar and vector product of two vectors with properties, Motion in a plane, cases of uniform velocity and uniform acceleration. Projectile motion.



Unit-IV : Laws of Motion

Marks 7

Concept of force and inertia, Newton's first law of motion, Momentum and Newton's second law of motion, impulse, Newton's Third Law of Motion. Law of conservation of linear momentum and its applications, Equilibrium of concurrent forces.

Friction, static and kinetic friction, laws of friction, rolling friction. Dynamics of uniform circular motion, centripetal force, examples of circular motion (vehicle on level circular road, vehicle on banked road).

Unit-V : Work, Energy and Power

Marks 6

Concept of scalar product of vectors, Work done by a constant force and a variable force, Kinetic Energy, Work energy theorem, Power.

Motion of potential energy, potential energy of spring, conservative forces, conservation of mechanical energy (K. E. and P. E's), non conservative forces, elastic and inelastic collision in one and two dimensions.

Unit-VI : Motion of system of particles and Rigid body.

Marks 6

Centre of mass of a two particle system, momentum, conservation and centre of mass motion, centre of mass of a rigid body, centre of mass of circular ring, disc, rod and sphere.

Concept of vector product of vectors: Moment of a force, torque, angular momentum, conservation of angular momentum with some examples.

Equilibrium of rigid bodies, rigid body rotation and equations of rotational motion, Comparison of linear and rotational motions, moment of inertia, radius of gyration.

Values of moment of inertia for simple geometrical objects (no derivation), statement of parallel and perpendicular axes theorem and their applications.

Unit VII: Gravitation

Marks 6

Kepler's laws of planetary motion, The universal law of gravitation. Acceleration due to gravity and its variation with altitude, depth and shape, Gravitational potential, gravitational



potential energy, escape velocity, orbital velocity of a satellite, geo-stationary satellite. Inertial and gravitational mass.

Unit VIII: Properties of Bulk matter

7 Marks

Elastic behaviour, stress-strain relationship, Hooke's law, young's modulus, bulk modulus, shear modulus of rigidity.

Pressure due to fluid column, Pascal's law and its applications (hydraulic lift and hydraulic brakes). Effect of gravity on fluid pressure.

Viscosity, stoke's law, terminal velocity, streamline and turbulent flow, Critical velocity, Reynold number, Bernoulli's theorem and its applications.

Surface energy and surface tension, angle of contact, applications of surface tension, ideas to drops, bubbles and capillary rise, action of detergents.

Heat, temperature, thermal expansion, specific heat, calorimetry, change of state-latent heat. Heat transfer-conduction, convection and radiation, thermal conductivity, Newton's law of cooling.

Unit IX : Thermodynamics

6 Marks

Thermal equilibrium and definition of temperature (Zeroth law of thermodynamics). Heat, work and internal energy, First law of thermodynamics. Second law of thermodynamics, reversible and irreversible processes. Heat engines and refrigerators (concept only).

Unit X : Behavior of perfect gas and Kinetic theory

6 Marks

Equation of state of perfect gas, work done on compressing a gas.

Kinetic theory of gases-assumptions, concept of pressure, expression for pressure exerted by a gas, Kinetic energy and temperature, rms speed of gas molecules, degrees of freedom, law of equipartition of energy (statement only) and application to specific heat capacities of gases, concept of mean free path, Avogadro's number.

Unit XI : Oscillation and waves

10 Marks

Periodic motion - period, frequency, displacement as a function of time. Periodic functions, simple harmonic motion (S.H.M) and its equation, phase, oscillation of a spring-restoring



force and force constant, energy in S.H.M-Kinetic and potential energies, simple pendulum- derivation of expression for its time period, free forced and damped oscillations (qualitative ideas only), resonance.

Wave motion - Longitudinal and transverse waves, speed of wave motion, Displacement relation for a progressive wave, Principle of super position of waves, reflection of waves, standing waves in strings and organ pipes, fundamental mode and harmonics. Beats, Doppler effect.

PRACTICALS

30 Marks

NOTE :- Every student is required to perform minimum of 10 experiments and 8 activities.

EXPERIMENTS :

1. Use of vernier calipers
 - i) To measure diameter of a small spherical/ cylindrical body.
 - ii) To measure internal diameter and depth of a given beaker/ calorimeter and hence find its volume.
2. Use of screw gauge.
 - i. To measure diameter of given wire.
 - ii. to measure thickness of a given sheet.
 - iii. to measure volume of an irregular lamina.
3. To determine radius of curvature of a given spherical surface by a spherometer.
4. To find the weight of a given body using parallelogram law of vectors.
5. Using a simple pendulum plot L-T graph hence find acceleration due to gravity (g).
6. To study the relation between force of limiting friction and normal relation force find coefficient of friction between a block and a horizontal pull of the earth and study in relationship with the angle of inclination by plotting a graph between force and $\sin \theta$.



ACTIVITIES/Project work

1. To make a paper scale of a given least count e.g. 0.2 cm, 0.5 cm.
2. To determine mass of given body using a meter scale by principle of moments.
3. To plot a graph for a given set of data, with proper choice of scales and error bars.
4. To measure the force of limiting friction for rolling of a roller on a horizontal plane.
5. To study the variation in range of jet of water with angle of projection.
6. To study dissipation of energy of a simple pendulum by plotting a graph between square of amplitude and time.
7. To study collision of two balls in two dimensions.

EXPERIMENTS

1. To determine young modulus of elasticity of the material of a given wire.
2. To find the force constant of a helical spring by plotting a graph between load and extension.
3. To determine the surface tension of water by capillary rise method.
4. To determine the coefficient of velocity of a given viscous fluid by measuring the terminal velocity of a given spherical body.
5. To find the speed of sound in air at room temperature using a resonance tube by two resonance position method.
6. To study relation between the length of a given wire and tension for constant frequency using sonometer.
7. To determine specific heat of a given solid and liquid, by method of mixtures.



ACTIVITIES/PROJECT WORK

1. To observe change of state and plot a cooling curve for melted wax.
2. To observe and explain the effect of heating on a bi-metallic strip.
3. To study the effect of detergent on surface tension by observing capillary rise.
4. To study the factors effecting the rate of loss of heat of a liquid.
5. To study the effect of nature of surface on emission and absorption of radiation.

Suggested Textbook: Textbook of Physics for class XI published by NCERT, New Delhi.



CHEMISTRY

Maximum Marks:100

Time 3 hrs.

Theory: 70 Marks

Practical: 30 Marks

UNIT-I: SOME BASIC CONCEPTS OF CHEMISTRY 05 Marks

General Introduction: Importance of studying chemistry, Historical approach to particulate nature of matter, Laws of Chemical combination (numerical), Dalton's Atomic Theory, Concept of elements, atoms & molecules. Atomic and molecular masses; Mole concept and molar mass, percentage composition, empirical and molecular formula; chemical reactions, stoichiometry and calculation based on stoichiometry.

Unit-II: STRUCTURE OF ATOM 05 Marks

Discovery of electron, proton and neutron, atomic number, isotopes and isobars. Thompson's model and its limitations, Rutherford's model and its limitations. Bohr's model & its limitations, concept of shells and sub-shells. Dual nature of matter and light, de-Broglie's relationship. Heisenberg's uncertainty principle, concept of orbitals, quantum numbers, shapes of s, p and d- orbitals. Rules for filling electrons in orbitals- Aufbau's principle, Pauli's exclusion principle and Hund's rule . Electronic configuration of atoms, stability of half filled and completely filled orbitals.

Unit-III: CLASSIFICATION OF ELEMENT AND PERIODICITY IN PROPERTIES 05 Marks

Significance of classification, brief history of the development of periodic table , modern periodic law and the present form of the periodic table, periodic trends in properties of elements: atomic radii, ionic radii, inert gas radii, ionization enthalpy, electron gain enthalpy, electronegativity, valence.



Unit-IV: CHEMICAL BONDING AND MOLECULAR STRUCTURE

05 Marks

Valence electrons, ionic bond, covalent bond, bond parameters, Lewis structure, polar character of covalent bond, valence bond theory, resonance, geometry of covalent molecules, VSEPR theory, concept of hybridization involving s, p and d orbitals and shapes of some simple molecules, molecular orbital theory of homonuclear molecules (Qualitative idea only), hydrogen bond.

Unit-V: STATES OF MATTER: GASES AND LIQUIDS

06 Marks

Three states of matter: intermolecular interactions, type of bonding, melting and boiling points, role of gas laws in elucidating the concept of the molecule, Boyle's law, Gay-Lussac's law, Avogadro's law, ideal behavior, empirical derivation of gas equation, Avogadro's number, ideal gas equation, deviation of real gases from ideal behavior, Liquefaction of gases, critical temperature.

Liquid state-vapour pressure, surface tension, viscosity (Qualitative idea only, no mathematical derivation).

Unit-VI: THERMODYNAMICS

04 Marks

Concepts of system, types of systems, surrounding, work, heat, energy, intensive and extensive properties, state functions. First Law of Thermodynamics, internal energy, enthalpy, heat capacity, specific heat, molar heat capacity, measurement of ΔE and ΔH ; Hess's law of constant heat summation, enthalpy of bond dissociation, combustion, formation, atomization, sublimation, phase transition ionization and dilution.

Introduction of entropy as a state function, free energy change for spontaneous and non-spontaneous process and equilibrium.

Unit-VII: EQUILIBRIUM

05 Marks

Equilibrium in physical and chemical processes, dynamic nature of equilibrium, law of mass action, equilibrium constant, factors affecting equilibrium: Le-Chatelier's principle; ionic equilibrium- ionization of acids and bases, strong and weak electrolytes,



degree of ionization, Concept of pH. Hydrolysis of salts (elementary idea), buffer solutions, solubility product, common ion effect (with suitable examples).

Unit-VIII: REDOX REACTIONS

02 Marks

Concept of oxidation and reduction, redox reactions, oxidation number, balancing of chemical equations in redox reactions, applications of redox reactions.

Unit-IX: HYDROGEN

02 Marks

Position of hydrogen in periodic table, occurrence, isotopes, preparation, properties and uses of hydrogen, hydrides-ionic, covalent and interstitial. Physical and chemical properties of water; heavy water; hydrogen peroxide-preparation, reactions and structure, hydrogen as a fuel,

Unit-X: s-BLOCK ELEMENTS (ALKALI AND ALKALINE EARTH METALS)

06 Marks

Group 1 and Group 2 elements;

General introduction, electronic configuration, occurrence, uses, anomalous properties of the first elements in each group, diagonal relationship; trends in the variation of properties (such as ionization enthalpy, atomic and ionic radii). Trends in chemical reactivity with oxygen, hydrogen, water and halogens; uses.

Preparation and properties of some important compounds: Sodium carbonate, Sodium chloride, sodium hydroxide and sodium hydrogen carbonate. Biological importance of sodium and potassium; CaO, CaCO₃ and industrial uses of lime and limestone, biological importance of Mg and Ca.

Unit-XI: SOME p-BLOCK ELEMENTS

05 Marks

General introduction to p-Block Elements

Group 13 elements: General introduction, electronic configuration, occurrence,



variation of properties, oxidation states, trends in chemical reactivity, anomalous properties of the first element in group. Boron - physical and chemical properties; some important compounds: borax, boric acids, boron hydrides.

Aluminium: uses, reactions with acids and alkalis.

Group 14 elements: General introduction, electronic configuration, occurrence, anomalous properties of the first element in group, trends in physical properties, trends in chemical properties. Carbon - catenation, allotropic forms, physical and chemical properties, trends in chemical properties, uses of oxides of carbon, important compounds of silicon and their uses: silicon tetrachloride, silicones, silicates and zeolites.

Unit-XII: ORGANIC CHEMISTRY- SOME BASIC PRINCIPLES AND TECHNIQUES

09 Marks

General introduction to organic chemistry, methods of purification, qualitative and quantitative analysis, classification and IUPAC nomenclature of organic compounds.

Electronic displacement in a covalent bond: inductive effect, electromeric effect, resonance and hyper-conjugation. Homolytic and heterolytic fission of a covalent bond, free radicals, electrophiles, nucleophiles, carbocations and carbanions. Types of organic reactions.

Unit-XIII: HYDROCARBONS

09 Marks

Classification of hydrocarbons

Alkanes: Nomenclature, isomerism, conformations (ethane only), physical properties, chemical reactions including free radical mechanism of halogenation, combustion and pyrolysis.

Alkenes: Nomenclature, structure of double bond (ethene), geometrical isomerism, methods of preparation, physical properties, chemical reactions- addition of hydrogen, halogen, water, hydrogen halides (Markownikov's addition and peroxide effect), ozonolysis, oxidation, mechanism of electrophilic addition.

Alkynes: Nomenclature, structure of triple bond (ethyne), physical properties, methods of preparation, chemical reactions: acidic character of alkynes, addition reaction of-



hydrogen, halogens, hydrogen halides and water, **Aromatic hydrocarbons:** Introduction, IUPAC nomenclature; Benzene: resonance, aromaticity; chemical properties; mechanism of electrophilic substitution - nitration, sulphonation, halogenation, Friedel Craft's alkylation and acylation; directive influence of functional group in monosubstituted benzene.

Unit-XIV: ENVIRONMENTAL CHEMISTRY

02 Marks

Environmental pollutions: soil, water and air pollution, acid rain, effects of the depletion of ozone layer, Green house effect and global warming- pollution due to industrial wastes. Lake water pollution: sources of pollutants in lake water, sources of pollution in Dal lake, Wullar lake and Mansar lake in J&K Green chemistry as an alternative tool for reducing pollution, strategy for control of environmental pollution.

PRACTICALS

Marks: 30

Time: 3 Hrs.

A) Organic Preparations:

- i) Preparation of acetylene and study of its acidic character.
- ii) Preparation of Acetanilide
- iii) Preparation of p-Nitroacetanilide

B) Characterization and Purification of Chemical Substance:

- i) Determination of melting point of an organic compound (below 100°C)
- ii) Determination of boiling point of an organic liquid.
- iii) Crystallization involving impure sample of any one of the following: Alum, Copper sulfate, Benzoic acid.

C) Experiments Related to pH Change

Any one of the following experiments:

- i) Determination of pH of some solutions obtained from juices and solutions of



known and varied concentrations of acids, bases and salts using pH paper/ universal indicator.

- ii) Comparing the pH of solutions of strong and weak acid of same concentration.
- iii) Study the pH change in the titration of a strong acid with a strong base using universal indicator.
- iv) Study of pH change by common-ion effect in case of weak acids and weak bases.

D) Chemical Equilibrium:

One of the following experiments:

- i) Study the shift in equilibrium between ferric ions and thiocyanate ions by increasing/ decreasing the concentration of either ions.
- ii) Study the shift in equilibrium between $[\text{Co}(\text{H}_2\text{O})_6]$ and Cl^- ions by changing the concentration of either ions.

E) Quantitative Estimation:

- i) Setting of a chemical balance and preparation of a standard solution of oxalic acid.
- ii) Determination of strength of a given sodium hydroxide solution by titrating it against a standard solution of oxalic acid.
- iii) Preparation of standard solution of sodium carbonate.
- iv) Determination of strength of given solution of dilute hydrochloric acid by titrating it against a standard solution of sodium carbonate.

F) Qualitative Analysis

Determination of one cation and one anion in a given salt (insoluble salts to be excluded):

Cations: Pb^{2+} , Cu^{2+} , As^{3+} , Al^{3+} , Fe^{3+} , Mn^{2+} , Zn^{2+} , Ni^{2+} , Co^{2+} , Ca^{2+} , Sr^{2+} , Ba^{2+} , Mg^{2+} , NH_4^+

Anions: CO_3^{2-} , S^{2-} , SO_3^{2-} , SO_4^{2-} , NO_2^- , NO_3^- , Cl^- , Br^- , PO_4^{3-} , $\text{C}_2\text{O}_4^{2-}$, CH_3COO^-



G) PROJECT

Scientific investigation involving laboratory testing and collecting information from other sources.

- Determination of BOD/ COD of locally available water sample.
- Analysis of fruit and vegetable juices for their acidity.
- Preparation of a sample of soap from available oils (Groundnut /Coconut oil).
- To dye wool and cotton clothes with any marked available dye.
- Study of the the effect of acids and bases on the tensile strength of fibres.
- Silvering of mirrors
- Compare the contents of tannic/ caffeine in various samples of tea and hence their flavor.

Note: Collaboration to sought from nearby institutions with regard to the performing of practicals/project work.

Suggested Textbook: A textbook of Chemistry for class XI published by NCERT, New Delhi



BIOLOGY

Maximum Marks: 100

Time: 3hrs.

Theory: 70 Marks

Practical: 30 Marks

SECTION A: (Botany)

Marks: 35

Unit-I Diversity of Life

8 marks

Variety of Living organism Systematics, need, history and classification (Artificial, natural and Phylogenetic). **Biosystematics**, Binomial nomenclature, Two kingdom system, five kingdom system, their merits and demerits. (Detailed study of kingdom,; Monera Protista and fungi), status of some acellular organisms/Slime moulds like: viruses and viroids. Lichens, taxonomic aids i.e. Botanical garden, herbaria, museum & keys.

Unit-II Kingdom Plantae

9 marks

Salient features of various plant groups for identification and their classes (Algae, Bryophytes, Pteridophytes, Gymnosperms and angiosperms). **Morphology of flowering plants and their function**. Morphology of root, stem, leaves, inflorescence, flowers, fruits and seed. Description of flowering plants of families Fabaceae, Solanaceae and Liliaceae.

Unit-III Anatomy of flowering plants

8 Marks

Tissues and tissue system, Types of Tissues, Meristematic and Permanent and their classification and functions.

Anatomy of Dicot and Monocot Root, Stem and Leaves, Secondary Growth in Dicot stems and roots.



Plant Physiology:

Transport in plants: means of transport, (diffusion, facilitated diffusion, Passive symports and anti ports, Active transport)

Plant water relations: water potential, osmosis, plasmolysis, imbibition, long distance transport of water- apoplast, symplast, pathways ascent of sap, Root pressure theory and transpirational pull theory (cohesion - tension theory).

Tranpiration: types & significance, mechanism of opening and closing of stomata, guttation, Phloem transport, flow from source to sink, (mass flow hypothesis)

Unit IV Mineral Nutrition

10 Marks

Methods to study mineral requirement (Hydroponics). Essential mineral, elements criteria for essentiality of nutrients. Essential elements. Micro and Macro nutrients, their role and deficiency symptoms. Mechanism of absorption of elements, translocation of solutes, soil and reservoir of essential elements. **Nitrogen metabolism**, Nitrogen cycle- Biological nitrogen fixation, 'Photosynthesis, Historical background, site of photosynthesis. Various photosynthetic pigments, Mechanism, Light reaction including PS I, PS II and photophosphorylation (Cyclic and non-cyclic). Dark reaction or Biosynthetic phase, Calvin (C_3) cycle, C_4 cycle, factors effecting photosynthesis. Photorespiration.

Respiration:- Introduction mechanism- glycolysis, Kreb's cycle. Electron transport system, Aerobic and anaerobic respiration. Respiratory quotient.

Growth and Development:- Characteristics of plant growth, phases of growth, growth curve and its components- differentiation, dedifferentiation and redifferentiation, **Development**, sequence of developmental processes in a plant cell, **plant growth regulators**, discovery and physiological effects (Auxins, Gibberellins, cytokinins, ethylene and IBA, Photoperiodism and vernalisation.



SECTION B: (ZOOLOGY)

Marks: 35

Unit I Diversity in Living world.

8 Marks

- i) Characteristic features of living organisms.
- ii) Salient features of animals (non chordates upto phylum level, chordates upto class level), Animal kingdom
- iii) Zoological parks, Natural museums (with special reference to local Zoos/National Parks (Manda, Mahamaya, Dachigam, Hemis)

Unit-II Cell-Structure and Function

10 Marks

- i) **Cell-** Brief description of cell, Cell theory; Prokaryotic and eukaryotic cell, cell wall, cell membrane and cell organelles (Plastids, Mitochondria, Endoplasmic reticulum, Golgi bodies/dictyosomes, Ribosomes, Lysosomes, Nucleus, Vacuoles, Centrioles), Cilia and flagella, and nuclear organization.
- ii) **Cell Division:-** Cell cycle, Mitosis, Meiosis.
- iii) Basic chemical constituents of living bodies.
- iv) **Biomolecules:** Structure and functions of :- carbohydrates, proteins, lipids and nucleic acids, Metabolites (Primary and Secondary, metabolism (elementary idea)
- v) **Enzymes:** Types, Properties and Functions.

Unit-III Histology and Morphology

7 marks

- i) **Animal tissues:-** Epithelial, Connective, Muscular & Nervous, Organ and Organ system
- ii) **Elementary Knowledge of :-** Morphology and Anatomy of Frog, Earthworm & Cockroach.

Unit IV Human Physiology

10 Marks

- i) Digestion and Absorption
- ii) Breathing and Respiration
- iii) Body fluids and circulation
- iv) Excretory products and elimination



- v) Locomotion and Movement
- vi) Neural control and coordination
- vii) Chemical coordination and integration.

PRACTICALS

M.Marks- 30

Time: 3 hrs.

SECTION A: (BOTANY)

Marks: 15

1. Study of different parts of a Compound Microscope.
2. Study of specimens and identification with reasons- Bacteria, Oscillatoria, Spirogyra, Rhizopus, Mushroom, Yeast, Liverwort (Marchantia/Moss (Funaria), Pinus (Male & female cone), Lichens.
3. Study of different modifications in
 - a. Roots (Tap & Adventitious)
 - b. Stems (Herbaceous & Woody)
 - c. Leaves (Leaf arrangement, shape, venation, simple & Compound leaves)
4. Description of 3 locally available flowers from the families- Fabaceae, Solanaceae and Liliaceae (1 from each family)
5. Study of plant tissues from permanent slides (Parenchyma, Collenchyma, Sclerenchyma, Xylem and Phloem)
6. Study of T.S. of Dicot & Monocot Root, Stem and leaf from permanent slides.
7. Study of osmosis by Potato osmoscope.
8. Study of Plasmolysis in epidermal peels(e.g. Rhoeo leaves)
9. Study of distribution of stomato in upper and lower surface of leaves
10. To make comparative study of the rates of transpiration in upper and lower surface of leaves by cobalt chloride method
11. Study of imbibition in seeds/ raisins
12. Observation and comment on the experimental set up on phototropism.
13. To separate plant pigments through paper chromatography.



SECTION-B (ZOOLOGY)

Marks: 15

1. Study and handling of compound Microscope.
2. Study of salient features of specimen and identification with reasons; Amoeba, Paramecium, Hydra, Liver fluke, Ascaris, Leech, Earth worm, Honeybee, Snail, Star fish, Shark, Labeo, Frog, Lizard and Pigeon.
3. Study of preserved specimens of at least one representative of each group to understand co-relations between characteristics of organisms and systematic position.
4. Study of animal cell and its organelles with the help of chart/slides.
5. Study of Mitosis and Meiosis from prepared slides.
6. Preparation of temporary mounts of mammalian squamous epithelium striped muscles, fibres and mammalian blood film.
7. Study of different types of mammalian connective tissues, muscle fibres and nerve cells through prepared permanent slides.
8. Study of different systems with the help of charts/dissections-Earthworm, Cockroach.
9. Testing for the presence of carbohydrate and protein.
10. Preparation and study of human blood smear.

Project work:

1. Collection of animal specimen for school museum.
2. Visit to a zoological /National park and preparation of report.
3. Study of cyclosis in Paramecium.
4. Study of Mitosis by using root tips of onion.
5. Study of Meiosis from flower buds.
6. Study of external morphology of earthworm, cockroach and frog.

Textbook Suggested: A Textbook of Biology for class XI published by NCERT, New Delhi.



PHYSICS

Maximum Marks: 100

Theory: Marks 70

Practicals: Marks 30

Time: 3 hour

I. Electrostatics	08 marks
II. Current Electricity	07 marks
III. Magnetic effects of current and magnetism	08 marks
IV. Electro-magnetic induction and alternating currents	08 marks
V. Electro-magnetic waves	03 marks
VI. Optics	14 marks
VII. Dual nature of matter and radiation	04 marks
VIII. Atoms and Nuclei	06 marks
IX. Electronic devices	07 marks
X. Communication system	05 marks

Unit I : Electrostatics

Electric charges; conservation of charge, coulomb's law – force between two point charges, forces between multiple charges, superposition principle and continuous charge distribution.

Electric field, electric field due to point charge, electric field lines, and electric dipole. electric field due to dipole, Torque on a dipole in uniform electric field.

Electric flux, statement of Gauss's theorem and its application to find field due to infinitely long straight wire, uniformly charged infinite plane sheet and uniformly charged thin spherical shell (field inside and outside).

Electric potential, potential difference, electric potential due to point charge, a dipole and system of charges; equipotential surfaces, electric potential energy of a system of two point charges and of electric dipole in an electrostatic field.

Conductor and insulators, free charges and bound charges inside a conductor. Dielectrics and electric polarization, capacitors and capacitance, combination of capacitors in series and in parallel, capacitance of a parallel plate capacitor with and without dielectric medium between the plates, energy stored in a capacitor. Van de Graaff generator.

Unit-II : Current Electricity

Electric current, flow of electric charges in a metallic conductor, drift velocity, mobility and their relation with electric current. Ohm's law, electric resistance. V-I Characteristics, (linear, non-linear), electrical energy and power, electric resistivity and conductivity, carbon resistors, colour code for carbon resistors; Temperature dependence of resistance.

Internal resistance of a cell, potential difference and emf of a cell, combination of cells in series and in parallel. Elementary idea of secondary cells. Kirchoff's laws and their applications. Wheat stone bridge, meter bridge.



Potentiometer-principle and its application to measure potential difference and for comparing e.m.f. of two cells; measurement of internal resistance of a cell.

Unit-III : Magnetic Effects of Current and Magnetism

Concept of magnetic field, Oersted's experiment, Biot-Savart law and its application to current carrying circular loop. Ampere's law and its applications to infinite long straight wire, straight and toroidal solenoids.

Force on a moving charge in a uniform magnetic and electric fields. Cyclotron. Force on a current carrying conductor in a uniform magnetic field. Force between two parallel current carrying conductors-definition of ampere.

Torque experienced by a current loop in uniform magnetic field, moving coil galvanometer-its current sensitivity and conversion to ammeter and voltmeter.

Current loop as a magnetic dipole and its magnetic dipole moment. Magnetic dipole moment of a revolving electron. Magnetic field intensity due to a magnetic dipole (bar magnet) along its axis and perpendicular to its axis. Torque on a magnetic dipole (bar magnet) in uniform magnetic field, bar magnet as an equivalent solenoid, magnetic field lines, Earth's magnetic field and magnetic elements. Para, dia, and ferro-magnetic substances with examples. Electromagnets and factors affecting their strength, permanent magnets.

Unit IV : Electro-magnetic Induction and Alternating Currents

Electromagnetic induction, Faraday's laws, induced e.m.f. and current; Lenz's law, Eddy currents, self and mutual inductance.

Alternating currents, peak and rms value of alternating current/voltage. Reactance and impedance, LC oscillations (qualitative treatment only) & LCR circuits series, Resonance, power in A.C. circuits, wattless current, AC Generator and transformer.

Unit-V : Electro-magnetic Waves

Need for displacement current, Electro-magnetic waves and their characteristics (qualitative ideas only), transverse nature of electromagnetic waves.

Electromagnetic spectrum (radio-waves, micro-waves, infra-red, visible, ultraviolet, X-rays, gamma rays) including elementary facts about their uses.

Unit VI : Optics

Ray Optics - Reflection of light; spherical mirrors; mirror formula, Refraction of light- total internal reflection and its applications, optical fibres, refraction at spherical surfaces, lenses, thin lenses formula, lens-makers formula, Newton's relation: displacement method to find position of images (conjugate points), Magnification, power of lens, combination of thin lenses in contact. Combination of a lens and a mirror, Refraction and dispersion of light through a prism.

Scattering of light-blue colour of the sky and reddish appearance of the sun at sunrise and sunset. Elementary idea of Raman effect.

Optical instruments - Human eye, image formation and accommodation, correction of eye defects (myopia, hypermetropia, presbyopia and astigmatism) using lenses. Microscopes and astronomical telescopes (reflecting and refracting) and their magnifying powers.

Wave optics-wave front and Huygen's principle, reflection and refraction of plane wave at



a plane surface using wavefronts. Proofs of laws of reflection and refraction using Huygen's Principle, Interference, Young's double slit experiment and expression for fringe width, coherent sources and sustained interference of light.

Diffraction due to a single slit, width of central maximum. Resolving power of microscopes and astronomical telescopes. Polarization, plane polarized light, Brewster's law, uses of plane polarized light and polaroids.

Unit VII : Dual Nature of Matter and Radiation

Dual nature of radiation. Photoelectric effect, Hertz and Lenard's observations; Einstein's photoelectric equation- particle nature of light.

Matter waves, wave nature of particles, de-Broglie relation, Davisson- Germer experiment (experimental details should be omitted; only conclusion should be explained).

Unit VIII : Atomic Nuclei

Alpha-particle scattering experiment, Rutherford's model of atom, Bohr's Model of atom; energy levels, Hydrogen spectrum. Continuous and characteristics of X-rays. Composition and size of nucleus; atomic masses, isotopes, isobars, isotones, Radioactivity (alpha, beta and gamma) particles/ rays and their properties, Radioactive decay law, Mass – energy relation, mass defect, binding energy/nucleon and its variation with mass no., nuclear fission and nuclear fusion.

Unit IX : Electronic Devices

Energy bands in solids, conductors, insulators and semiconductors, semiconductor diode, I-V characteristics in forward and reverse bias, diode as a rectifier; I-V characteristics of LED, photo diode, solar cell and Zener diode; Zener diode as a voltage regulator, Junction transistors and its action; characteristics of a transistor, transistor as an amplifier (common emitter configuration and oscillator (common emitter). Logic gates (OR, AND, NOT), concept of NAND and NOR gates, Transistor as a switch.

Unit X : Communication System

Elements of communication system (block diagram only), Band width of signals (speech, T.V and digital data); bandwidth of transmission medium, propagation of electromagnetic waves in the atmosphere, sky and space wave propagation.

Need for modulation; Production and detection of an amplitude modulated wave.



Practicals : 30 marks

External: 20

Internal:10

Every student will perform at least 15 experiments (7 from section A & 8 from section B). The activities mentioned here should be for the purpose of demonstration. One project of three marks is to be carried out by the students.

Evaluation Scheme for Practical Examination:

- One experiment from each of the two sections = 10 marks
- One activity from each of the two sections (2 activities in total) = 2+2= 04 marks
- Record of one Investigatory Project and viva based on Project = 02 marks
- Practical Record of experiments and activities = 02 marks
- Viva-voce on experiments and activity = 02 marks

Total Marks = 20

Section – A

Experiments:

1. To determine resistance per cm. of a given wire by plotting a graph of pot. difference vs. current (Ohm's law).
2. To find resistance of a given wire using meter bridge and hence determine the specific resistance of its material.
3. To verify the laws of combination (series/parallel) of resistance using a metre bridge.
4. To compare the e.m.f of two given primary cells using potentiometer.
5. To determine internal resistance of a given primary cell using potentiometer.
6. To determine resistance of a galvanometer by using half deflection method and also find its figure of merit.
7. To convert the given galvanometer (of known resistance and figure of merit) into an ammeter and voltmeter of desired range and to verify the same.
8. To find the frequency of the a.c. mains with a Sonometer

Activities:

1. To measure the resistance and impedance of an inductor with or without iron core.
2. To measure resistance voltage (AC/DC), current (AC) and check continuity of a given circuit using multi metre.
3. To assemble a household circuit comprising three bulbs, three (on/off) switches, a fuse and a power source.
4. To study the variation in potential drop with length of a wire for a steady current.
5. To draw the diagram of a given open circuit comprising at least a battery, rheostat, key;



ammeter, voltmeter **Make the components that are not connected in proper order and correct the circuit and also circuit diagram.**

Section - B

Experiment:

1. To find the focal length of a convex mirror, using a convex lens.
2. To find the focal length of a concave lens using a convex lens.
3. To find the value of v for different values of u in case of a concave mirror and also to find its focal length.
4. To find the focal length of a convex lens by plotting a graph between u and v or between $1/u$ and $1/v$.
5. To determine angle of minimum deviation (d_m) for a given prism by plotting a graph between angle of incidence and angle of deviation (d).
6. To determine refractive index of a glass slab using a travelling microscope.
7. To find refractive index of a liquid using I) concave mirror II) convex lens and plane mirror.
8. To draw the characteristics of a common-emitter npn or pnp transistor and to find out the values of current and voltage gains.
9. To draw the I-V characteristics curve of a p-n junction in forward bias and reverse bias.
10. To draw the characteristic curve of a zener diode and to determine its reverse break down voltage.

Activities:

1. To study effect of intensity of light by varying distance of the source on an L.D.R.
2. To identify a diode, a LED, a transistor, and IC, a resistor and a capacitor from mixed collection of such items.
3. Use of multimeter to i) identify base of transistor ii). Distinguish between npn and pnp-transistors iii) see the unidirectional flow of current in case of a diode and an LED. iv) Check whether a given electronic component (e.g. diode, transistor or IC) is in working order.
4. To observe refraction and lateral deviation of a beam of light incident obliquely on a glass slab.
5. To observe polarization of light using two polaroids.
6. To observe diffraction of light due to a thin slit.
7. To study the size and nature of the image formed by i) convex lens, ii) concave mirror, on a screen by using a candle and screen for different distances of the candle from the lens/mirror.
8. To obtain a lens combination with the specified focal length by using two lenses from the given set of lenses.



Investigatory Projects:

1. To investigate whether the energy of a simple pendulum is conserved.
2. To determine the radius of gyration about the centre of mass of a scale used as a bar pendulum.
3. To investigate changes in the velocity of a body under the action of a constant force and determine its acceleration.
4. To compare effectiveness of different materials as absorbers of sound or heat.
5. To determine the wave length of laser beam by diffraction.
6. To study various factors on which the internal resistance, emf of a cell depends.
7. To construct a time switch and study dependence of its time constant on various factors.
8. To study infrared radiations emitted by different sources using photo-transistor.
9. To compare effectiveness of different materials and insulators.
10. To design an automatic traffic signal system using suitable combination of logic gates.
11. To study luminosity of various electric lamps of different powers and make.
12. To compare the Young's modulus of elasticity of different specimens of rubber and also draw their elastic hysteresis curve.

Book Suggested: A Textbook of Physics for class XII published by NCERT, New Delhi.



CHEMISTRY

Maximum Marks: 100

Theory: Marks 70

Time: 3 hour

Practicals: Marks 30

Unit I	Solid State	4 marks
Unit II	Solutions	5 marks
Unit III	Electrochemistry	5 marks
Unit IV	Chemical Kinetics	5 marks
Unit V	Surface Chemistry	4 marks
Unit VI	General Principles and Processes of Isolation of Elements	3 marks
Unit VII	p-Block Elements	8 marks
Unit VIII	d- and f- Block Elements	5 marks
Unit IX	Coordination Compounds	3 marks
Unit X	Haloalkanes and Haloarenes	4 marks
Unit XI	Alcohols, Phenols and Ethers	4 marks
Unit XII	Aldehydes, Ketones and Carboxylic Acids	6 marks
Unit XIII	Organic Compounds containing Nitrogen	4 marks
Unit XIV	Biomolecules	4 marks
Unit XV	Polymers	3 marks
Unit XVI	Chemistry in Everyday Life	3 marks

Unit-I: SOLID STATE

Classification of solids based on different binding forces: molecular, ionic, covalent and metallic solids, amorphous solids and crystalline solids (elementary idea only), unit cell in two dimensional & three dimensional lattices, packing efficiency, calculation of density of unit cell, packing in solids, voids, number of atoms per unit cell in a cubic unit cell, point defects. Properties of solids (electrical, magnetic & dielectric), Band theory of metals, conductors, semi-conductors and insulators and n & p type semiconductors.

Unit-II: SOLUTIONS

Types of solutions, expression of concentration of solutions of solids in liquids, solubility of gases in liquids, solid solutions, colligative properties: relative lowering of vapor pressure of a solution, Raoult's law, elevation of boiling point, depression in freezing point temperature and osmotic pressure), determination of molecular masses using colligative properties. Abnormal molecular mass, van't Hoff factor and calculations involving it..

Unit-III: ELECTROCHEMISTRY

Redox reactions, conductance in electrolytic solutions, specific conductivity, molar



conductivity, variation of conductivity with concentration, Kohlrausch's law and its applications Electrolysis and laws of electrolysis (elementary idea), dry cell- electrolytic cells and galvanic cells; lead accumulator, emf of a cell, standard electrode potential, Nernst equation and its application to chemical cells, relation between Gibb's energy change and emf of a cell, fuel cells, corrosion

Unit-IV: CHEMICAL KINETICS

Rate of reaction (average and instantaneous rate of a reaction), factors affecting rate of reactions: (concentration, temperature, catalyst), rate law, specific rate constant and order, molecularity of a reaction, integrated rate expression of zero and first order reactions and their derivations, half life period. Concept of collision theory (elementary idea, no mathematical derivation), Activation energy, Arrhenius equation.

Unit-V: SURFACE CHEMISTRY

Adsorption- physical and chemical adsorption, factors affecting adsorption of gases on solids; Catalysis: homogeneous and heterogeneous, activity & selectivity. Enzyme catalysis, Colloidal state: distinction between true solution, colloids and suspensions. Types of colloids- lyophilic and lyophobic, multimolecular, macromolecular and associated colloids (micelles), properties of colloids: Tyndall effect, Brownian movement, Electrophoresis, Coagulation, Emulsions-types of emulsions. Elementary idea about nanomaterials.

Unit-VI: GENERAL PRINCIPLES AND PROCESSES OF ISOLATION OF ELEMENTS

Principles and methods of extraction: concentration, oxidation, reduction, electrolytic method & refining; occurrence & principles of extraction of aluminium, copper, zinc and iron.

Unit- VII: p-BLOCK ELEMENTS

Group 15 Elements: General introduction, electronic configuration, occurrence, oxidation states, trends in physical and chemical properties; nitrogen: preparation, properties & uses. Compounds of nitrogen: preparation & properties of ammonia and nitric acid, oxides of nitrogen (structure only), Phosphorus – allotropic forms; compounds of phosphorus: preparation & properties of phosphine, halides (PCl_3 , PCl_5) and oxo- acids (elementary idea only).

Group 16 Elements: General introduction, electronic configuration, occurrence, oxidation states, trends in physical and chemical properties; dioxygen: preparation, properties & uses. Classification of oxides; ozone. Sulphur- allotropic forms; compounds of sulphur: preparation, properties & uses of SO_2 and Sulphuric acid: industrial process of manufacture, properties and uses, other oxides and oxoacids of sulphur (structures only).

Group 17 Elements: General introduction, electronic configuration, oxidation states, trends in physical and chemical properties; compounds of halogens-preparation, properties and uses of Chlorine and hydrochloric acid, interhalogen compounds, oxoacids of halogens (structures only)

Group 18 Elements: General introduction, electronic configuration, occurrence, trends in physical & chemical properties & Uses.

Unit- VIII: d and f –BLOCK ELEMENTS

General introduction, electronic configuration, occurrence and characteristics of the transition metals, general trends in properties of first row transition metals (metallic character, IE, electrode



potential, oxidation state, ionic radii, catalytic properties, colored ions, complex formation, magnetic properties, interstitial compounds, alloy formation). Preparation and properties of $K_2Cr_2O_7$ and $KMnO_4$.

Lanthanides: electronic configuration, oxidation state, chemical reactivity and lanthanide contraction and its consequences.

Actinides- electronic configuration, oxidation states and comparison with lanthanoids.

Unit- IX: CO-ORDINATION COMPOUNDS

Co-ordination compounds: Introduction, ligands, co-ordination number, color, magnetic properties and shapes, IUPAC nomenclature of mononuclear co-ordination compounds. Bonding (Werner's theory, VBT and CFT); structural and stereoisomerisms, importance of coordination compounds in qualitative inclusion of analysis, extraction of metals and biological systems.

Unit-X: HALOALKANES AND HALOARENES

Haloalkanes: Nomenclature, nature of C-X bond, physical & chemical properties, mechanism of substitution reactions. Stability of carbocations, R-S and d-l configurations.

Haloarenes: Nature of C-X bond, substitution reactions (directive influence of halogens for monosubstituted compounds only), Stability of carbocations, R-S and D-L configurations

Uses and environmental effects of- dichloromethane, trichloromethane, tetrachloromethane, iodoform, freon, and DDT.

Unit- XI: ALCOHOLS, PHENOLS AND ETHERS

Alcohols: Nomenclature, methods of preparation, physical & chemical properties (of primary alcohols only), identification of primary, secondary & tertiary alcohols; mechanism of dehydration of alcohols, uses, some important compounds – methanol and ethanol.

Phenols: Nomenclature, methods of preparation, physical & chemical properties, acidic nature of phenol, electrophilic substitution reactions, uses of phenols.

Ethers: Nomenclature, methods of preparation, physical & chemical properties and uses.

UNIT- XII: ALDEHYDES, KETONES AND CARBOXYLIC ACIDS

Aldehydes and Ketones: Nomenclature, nature of carbonyl group, methods of preparation, physical & chemical properties & mechanism of nucleophilic addition reaction to C = O group, reactivity of alpha hydrogen in aldehydes, uses.

Carboxylic Acids: Nomenclature, acidic nature, methods of preparation, physical & chemical properties and uses

UNIT- XIII: ORGANIC COMPOUNDS CONTAINING NITROGEN

Amines: Nomenclature, classification, structure, methods of preparation, physical & chemical properties, uses, identification of primary, secondary & tertiary amines.

Cyanides and Isocyanides: Structures of cyanide and isocyanide groups, nomenclature, preparation, physical properties and chemical reactions.



Diazonium Salts: Preparation and chemical reactions (mechanism of diazotization), and importance in synthetic organic chemistry.

UNIT- XIV: BIOMOLECULES

Carbohydrates: Classification (aldoses and ketoses), monosaccharides: Glucose, fructose: structure, preparation and chemical reactions; oligosaccharides (sucrose, lactose & maltose) Polysaccharides: (starch, cellulose and glycogen); importance.

Proteins: Elementary idea of amino acids: peptide bond, polypeptides and primary, secondary, tertiary and quaternary structure of proteins (Qualitative idea only). denaturation of proteins; enzymes, lipids & hormones, their classification & functions.

Nucleic Acids: DNA and RNA (purines and pyrimidines, nucleosides, nucleotides and fragments up to four nucleotides).

Vitamins: Classification and functions, sources and deficiency diseases.

UNIT- XIV: POLYMERS

Natural & synthetic polymers, methods of polymerization (addition and condensation), copolymerization, and some important polymers: natural and synthetic like polythene, nylon, Bakelite, polyesters and rubber. Biodegradable and non- biodegradable polymers.

Unit-XVI: CHEMISTRY IN EVERYDAY LIFE

- i) **Chemicals in medicine and health care-** analgesics, tranquillizers, antiseptics, disinfectants, antimicrobials, anti-fertility drugs, anti-histamines, antibiotics, antacids.
- ii) **Chemicals in food-** preservatives, artificial sweetening agents.
- iii) **Cleansing agents** – Soaps and detergents, cleansing action.

PRACTICALS

External: 20

Internal:10

Evaluation Scheme for Practical Examination: (External)

- Volumetric analysis	=	06 marks
- Salt Analysis	=	06 marks
- Content based experiment	=	04 marks
- Class record, Project work and viva	=	04 marks
		Total = 20 marks

A. SURFACE CHEMISTRY

- i) Preparation of one lyophilic and one lyophobic sol
Lyophilic sol-starch, egg albumin and gum
Lyophobic sol-aluminium hydroxide, ferric hydroxide, arsenious sulphide.
- ii) Study of the role of emulsifying agents in stabilizing the emulsion of different oils.



B. CHEMICAL KINETICS

- i) Effect of concentration and temperature on the rate of reaction between sodium thiosulphate and hydrochloric acid.
- ii) Study of reaction rates of any one of the following:
 - a) Reaction of iodide ions with hydrogen peroxide at room temperature using different concentration of iodide ions.
 - b) Reaction between potassium iodate (KIO_3) and sodium sulphite (Na_2SO_3) using starch solution as indicator (clock reaction).

C. THERMOCHEMISTRY

Any one of the following experiments:

- i) Enthalpy of dissolution of CuSO_4 or KNO_3 .
- ii) Enthalpy of neutralization of strong acid (HCl) and strong base (NaOH).
- iii) Determination of enthalpy change during interaction (Hydrogen bond formation) between acetone & chloroform.

D. ELECTRO CHEMISTRY

- i) Variation of cell potential in $\text{Zn}/\text{Zn}^{2+} // \text{Cu}^{2+}/\text{Cu}$ with change in concentration of electrolytes (CuSO_4 or ZnSO_4) at room temperature.

E. CHROMATOGRAPHY

- i) Separation of pigments from extracts of leaves and flowers by paper chromatography and determination of R_f values.
- ii) Separation of constituents present in an inorganic mixture containing two cations only (constituents having wide difference in R_f values to be provided)

F. PREPARATION OF INORGANIC COMPOUNDS

- i) Preparation of double salt of ferrous ammonium sulphate or potash alum.
- ii) Preparation of potassium ferric oxalate.

G. TEST FOR THE FUNCTIONAL GROUPS PRESENT IN ORGANIC COMPOUNDS

Unsaturation, alcoholic, phenolic, aldehydic, ketonic, carboxylic and amino (primary) groups.

H. Study of carbohydrates, fats and proteins in pure form and detection of their presence in given foodstuffs.

I. Determination of concentration/ molarity of KMnO_4 solution by titrating it against a standard solution of:

- i) oxalic acid
- ii) ferrous ammonium sulphate.



J. Qualitative Analysis

Determination of one cation and one anion in a given salt (insoluble salts to be excluded):

Cations: Pb^{2+} , Cu^{2+} , As^{3+} , Al^{3+} , Fe^{3+} , Mn^{2+} , Zn^{2+} , Ni^{2+} , Co^{2+} , Ca^{2+} , Sr^{2+} , Ba^{2+} , Mg^{2+} , NH_4^+

Anions: CO_3^{2-} , S^{2-} , SO_3^{2-} , SO_4^{2-} , NO_2^- , NO_3^- , Cl^- , Br^- , I^- , PO_4^{3-} , $\text{C}_2\text{O}_4^{2-}$, CH_3COO^-

PROJECT WORK

Wherever feasible may include

1) Model Preparation

2) Investigatory Project

- To prepare rayon thread from filter paper by cupra ammonium process.
- Determine the oxalate content of Guava fruits at different stages of ripening.
- Study of insecticides and pesticides in various fruits and vegetables.
- To determine the amount of casein present in different samples of milk from different sources.
- Preparation of soyabean milk and its comparison with natural milk.
- To determine the presence of adulterants in common foods such as sugar, butter, oil, red chilly paper, turmeric powder, rice.
- Prevention of rusting of iron by using cathode protection method.

3) Science Exhibits.

4) Participation in Science Fairs

Book Suggested: Textbook of Chemistry for class XII published by NCERT, New Delhi



BIOLOGY

Maximum Marks: 100

Theory: Marks 70

Practicals: Marks 30

Time: 3 hour

SECTION A (Botany)

Marks: 35

Unit-I: Reproduction in Flowering Plants

Marks:07

Asexual Reproduction: Vegetative propagation in plants, micropropagation.

Sexual Reproduction: Flower structure, Development of male & female gametophytes. Pollination: types, agencies & examples, Out breeding devices. Pollen- Pistil interaction, Double fertilization, Post fertilization events, Development of endosperm, embryo, seed and fruit. Special modes: apomixis and polyembryony, significance of seed & fruit formation.

Unit-II: Genetics

Marks 09

- Heredity and variation
- Mendelian inheritance, Deviations from Mendelism: incomplete dominance, co-dominance, Multiple alleles, Pleiotropy, Chromosomal theory of inheritance, Elementary idea of polygenic inheritance, Chromosomes & genes,
- Search for genetic material & DNA as genetic material: Structure of DNA & RNA, DNA packaging, DNA Replication (Semiconservative), Central dogma, Protein Biosynthesis: Transcription, translation, genetic code, Gene expression and regulation (lac-operon).

Unit-III: Biology and Human welfare

Marks: 07

- **Plant breeding:** Introduction, steps in plant breeding and application of plant breeding, and single cell protein, Biofortification.
- **Tissue culture:** Cellular totipotency, technique and application of tissue culture
- **Microbes in Human Welfare:** in household food processing, industrial production, sewage treatment, Production of energy (Biogas), biocontrol agent (Biopesticides) & Biofertilizers.
- Genetically Modified organism- Bt crops
- Biopiracy and patents.

Unit- IV: Ecology and Environment

Marks: 12

Meaning of ecology, environment, habitat and niche: Organisms and environment.

Population and ecological adaptations: Population Interactions-mutualism, competition, predation, parasitism. Population attributes-growth, birth rate and death rate, age distribution.

Ecosystems: Patterns, Components, energy flow, nutrient cycling (carbon and phosphorus), decomposition and productivity. Pyramids of number, biomass, energy. Ecological succession. Ecological Services: Carbon fixation, Pollination, Oxygen release.

Biodiversity and its conservation: Threats to, and need for biodiversity conservation. Hotspots, endangered organisms, extinction, Red Data Book. Biodiversity conservation-biosphere reserves, national parks and sanctuaries.



Environmental Issues: Air and water pollution and their control, solid waste management, agrochemicals and their effects, Radioactive waste management, Green house effect and global warming, Ozone depletion in stratosphere, Deforestation, Any three case studies as success stories addressing environmental issues.

SECTION B (Zoology)

35 Marks

Unit-I : Reproduction

Marks 11

- i) **Asexual Reproduction:** Uniparental, modes: binary fission, sporulation, budding, gemmule, fragmentation, regeneration.
- ii) **Human Reproduction-** Male and female reproductive systems, Microscopic anatomy of testis & ovary; Gametogenesis (spermatogenesis & oogenesis. Menstrual cycle), Fertilization, embryo development upto blastocyst formation, implantation; Pregnancy and placenta formation (elementary idea), Parturition (elementary idea) and Lactation (elementary idea).
- iii) **Reproductive Health:** Need for reproductive health & prevention of Sexually Transmitted Diseases (STD), Birth control- need & methods, Contraception and Medical Termination of Pregnancy (MTP), Amniocentesis, Infertility & assisted reproductive technologies: IVF, ZIFT, GIFT (Elementary idea for general awareness).

Unit-II: Genetics and Evolution

Marks 12

- Sex determination in humans, birds and honeybee.
- Inheritance pattern of Hemophilia and Color blindness in human beings.
- Mendelian Disorders in humans: Chromosomal disorders in humans, Down's syndrome, Turner's & Klinefelter's syndromes.
- Genome and Human Genome project.
- DNA fingerprinting.
- Origin of life: Theories & evidences with special reference to Darwin & Modern Synthetic theory of evolution, Hardy – Weinberg's principal. Adaptive radiation.
- Origin and evolution of Man.

Unit-III : Biology and Human Welfare

Marks 07

- **Health and Disease:** Basic concepts of immunology, vaccines; pathogens, parasites causing human diseases (Typhoid, Hepatitis, Malaria, Filariasis, Ascariasis, Common Cold, Amoebiasis, Ring Worm); Cancer, HIV and AIDS.
- **Insects & human welfare:** Silk, honey, lac.
- Adolescence, drug & alcohol abuse.
- Poultry, Dairy Farming

Unit IV: Biotechnology and its Application

Marks 05

- i) Genetic Engineering (Recombinant DNA technology), cloning
- ii) Applications in Health: Human insulin & vaccine production, gene therapy
- iii) Biosafety issues.



Practicals
External: 20

Time: 3 Hours
Internal:10

Marks: 30

Botany based Practicals: 15 Marks

- i) Internal assessment: 05 marks
- ii) External assessment: 10 marks

Zoology based Practicals: 15 Marks

- i) Internal assessment: 05 marks
- ii) External assessment: 10 marks

List of Experiments

1. Study pollen germination on a slide.
2. Collect and study soil from at least two different sites and study them for texture, moisture content, pH and water holding capacity of soil. Correlate with the kinds of plants found in them.
3. Collect water from two different water bodies around you and study them for pH, clarity and presence of any living organisms.
4. Study the presence of suspended particulate matter in air at the two widely different sites.
5. Study of plant population density by quadrat method.
6. Study of plant population frequency by quadrat method.
7. Prepare a temporary mount of onion root tip to study mitosis.
8. To study the effect of the different temperatures and three different pH on the activity of salivary amylase on starch.

Study/observation of the following (Spotting)

1. Flowers adapted to pollination by different agencies (wind, insect)
2. Pollen germination on stigma through a permanent slide.
3. Identification of stages of gamete development i.e. T.S. testis and T.S. ovary through permanent slides. (from any mammal)
4. Meiosis in onion bud cell or grasshopper testis through permanent slides.
5. T.S. of blastula through permanent slides.
6. Mendelian inheritance using seeds of different color / size of any plant.
7. Prepared pedigree charts of genetic traits such as rolling of tongue, blood groups, widow's peak, and color blindness.
8. Exercise on controlled pollination-Emasculation, tagging and bagging.
9. Identification of common disease causing organisms like Ascaris, Entamoeba, Plasmodium, Ringworm through permanent slides or specimens. Comment on symptoms of diseases that they cause.
10. Two plants and two animals found in xerophytic conditions. Comment upon their morphological adaptations.
11. Plants and animals found in aquatic conditions. Comment upon their morphological adaptations.

Book Prescribed: A Textbook of Biology for class XII published by NCERT, New Delhi.